Page 379 #2,4,7,9,12,26,32,42,87,94,95 Page 427 #8

2. a. Na^+ has a charge, Ne does not
   b. Cl_2 & Cl are both neutral, Cl^- has a charge

4. a. $[Ne]3s^23p^6$ K^+
   b. $[Ne]3s^23p^6$ S^{2-}
   c. $[He]2s^22p^6$ Al^{3+}
   d. $[Ne]3s^23p^6$ Cl^-
   e. $[He]2s^22p^6$ Mg^{2+}
   f. $[He]2s^22p^6$ N^{3-}

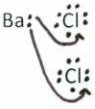
7. 

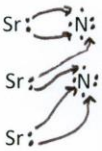
9. a. N b. Cl c. F d. O

12. a. 1 b. 4 c. 2 d. 1 e. 3 f. 1

26. a.  CaO

   b.  KBr

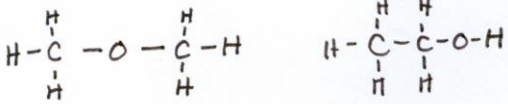
   c.  $BaCl_2$

   d.  Sr_3N_2

32. Lattice Energy = -2809 kJ
   Li_2O You will need a 1st and 2nd Electron Affinity

42. a. F-F, Cl-F, Br-F, I-F, H-F

   b. H-H, H-I, H-Br, H-Cl, H-F

87. 

94. -15kJ MgCl

95. -616 kJ $MgCl_2$
   Compare the answers for 94 & 95. The much larger value in 95 indicates $MgCl_2$ is the compound that is most likely to form and will be the most stable.

Page 427

8. no, no